THE DETERMINATION OF SERUM IRON AFTER THE INTRAVENOUS INJECTION OF IRON-DEXTRAN

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A rapid technique for the determination of iron in serum, after the intravenous injection of iron-dextran, is presented.

Dithionite is used for the reduction of ferric to ferrous iron.

The use of a detergent solution enables protein precipitation to be omitted.

The iron reagent, 4:7-diphenyl-1:10-phenanthroline, and its ferrous complex are both soluble in the dilute detergent solution.

Lipaeemic serum can be used without affecting the results.

The recovery of added ferric iron, as iron-dextran, is quantitative.

The method is presented.

Method

Apparatus and Reagents.—All apparatus should be rendered iron-free by any of the conventional methods. Thereafter the test-tubes simply require rinsing in distilled water after each batch of determinations. Although the “teepol” used was ordinary commercial grade as supplied for cleansing purposes it was surprisingly free from iron and required no preliminary treatment. Teepol powder (“teepol XL”) can also be obtained, but this contains considerably more iron than the liquid concentrate and was not used in this study. The distilled water used throughout was rendered iron-free by passage through an “elgital” ion-exchange resin. The reagent blank was of the order of 0.2 μg. Fe/determination, but about half of this “iron” was due to the colour of the iron reagent.

Iron Reagent.—Dissolve 5 mg. 4:7-diphenyl-1:10-phenanthroline in 10 ml. liquid “teepol” concentrate (27% active ingredient) by warming on a hotplate and breaking down any lumps with a glass rod. After cooling make the volume to 100 ml. with distilled water. When larger volumes of iron reagent are required the preparation of a stock solution containing 50 mg. bathophenanthroline/100 ml. “teepol” concentrate is recommended. The working reagent is obtained by diluting this stock solution 1 in 10 with water. The reagent appears to keep indefinitely.

Reducing Agent.—Solid sodium hydrosulphite (dithionite) Na2S2O4·H2O, which should be of the highest available purity.

Standard Iron Solution (1 mg. Fe/ml.).—Dissolve 702 mg. A.R. grade ferrous ammonium sulphate ((NH4)2SO4·FeSO4·6H2O) in N/10 sulphuric acid and make the volume to exactly 100 ml. with the acid.
**DETERMINATION OF SERUM IRON AFTER IRON-DEXTRAN INJECTION**

**Procedure.**—Treat both blank and test sera (or heparinized plasma) as follows:

 Deliver 0.2 ml. serum into a 15 ml. “pyrex” test-tube containing 3.8 ml. iron reagent and 10–20 mg. dithionite.

 Mix by lateral shaking and place in a water-bath at 90–95° C. for three minutes.

 Remove and cool to room temperature.

 Read the optical density of the test solution against the blank in the spectrophotometer at 530 mμ.

 **Standards.**—Prepare working standards by diluting the standard iron solution (1 mg. Fe/ml.) 3 in 200 and 3 in 100 with distilled water to give solutions containing 15 and 30 μg. Fe/ml. Treat 0.2 ml. of these diluted standards (S₁ and S₂) exactly as described for serum or plasma and read the optical densities of the resulting solutions in the spectrophotometer against a reagent blank. As there is complete agreement with Beer’s law the optical density of S₁ should be exactly twice that of S₂. Using a “unicam” SP. 600 spectrophotometer at 530 mμ and with 10 mm. cells, an optical density of 0.092/μg. Fe was obtained.

 **Calculation.**—This is as follows:

 \[ \text{Iron content of test} = \frac{\text{O.D. of test} \times \text{O.D. of test}}{\text{O.D. of S₁} \times 15, \text{or O.D. of S₂} \times 30 \text{ μg./ml.}} \]

 For instruments which require up to 5 ml. of final solution the volume of iron reagent may be increased to 4.8 ml.

 **Iron Injected.**—To determine the exact amount of iron injected, dilute the injection solution with distilled water to contain approximately 20 μg. Fe/ml. Treat 0.2 ml. of this solution exactly as described for serum or plasma and read in the spectrophotometer against a reagent blank to obtain μg. Fe/0.2 ml. solution.

 Then

 \[ \text{Iron injected per ml.} = \frac{\mu g. \text{ Fe/0.2 ml.} \times 5 \times \text{dilution factor}}{1000} \text{ mg.} \]

 The plasma volume can now be calculated in the usual way by dividing the μg. iron injected by the difference in iron content of the blank and test sera.

 \[ \text{Plasma volume} = \frac{\mu g. \text{ iron injected}}{\text{iron content of test sample (μg./ml.)}} \text{ ml.} \]

 **Results**

 Recovery experiments were carried out by adding known amounts of iron, in the form of iron-dextran, to pooled serum. Analysis of the resulting sera by the proposed method showed that the recovery was quantitative (Table I).

 **Table I**

<table>
<thead>
<tr>
<th>Iron Added (μg. Fe/100 ml. Serum)</th>
<th>Iron Found (μg. Fe/100 ml. Serum)</th>
<th>% Recovery</th>
</tr>
</thead>
<tbody>
<tr>
<td>500</td>
<td>515</td>
<td>103</td>
</tr>
<tr>
<td>1,000</td>
<td>1,000</td>
<td>100</td>
</tr>
<tr>
<td>1,500</td>
<td>1,480</td>
<td>98.7</td>
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<td>2,000</td>
<td>2,050</td>
<td>102.5</td>
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<td>2,500</td>
<td>2,490</td>
<td>99.6</td>
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 **Discussion**

 Under the conditions of the test the bathophenanthroline reagent is specific for ferrous iron. Buffering of the reagent was considered unnecessary since the final solution is maintained at an acid pH by the addition of the dithionite; colour development is maximum over the pH range 3.5–7.0.

 If the iron content of the test serum is low the volume of serum used can be increased to 0.5 ml., 3.5 ml. of the iron reagent then being used. With an injection of 50 mg. iron and a plasma volume of 3 litres the optical density of the final test solution will be about 0.30 using 0.2 ml. serum.

 The reduction of ferric to ferrous iron with dithionite is complete within two minutes at room temperature, but heating at 90 to 95° C. is necessary to render the final solution optically clear. Due to a slight turbidity unheated serum solutions have a higher optical density than the corresponding heated solution, and this difference decreases slowly during the first hour after mixing. The use of dithionite reduces the time for a plasma volume estimation to little more than 30 minutes from the time of injection, if heparinized.

 **Table II**

<table>
<thead>
<tr>
<th>Optical Density</th>
<th>Serum Sample (μg. Fe/ml.)</th>
<th>Optical Density</th>
<th>Serum Sample (μg. Fe/ml.)</th>
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<tr>
<td>0.377</td>
<td>0.338</td>
<td>0.377</td>
<td>0.336</td>
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<tr>
<td>0.380</td>
<td>0.334</td>
<td>0.373</td>
<td>0.338</td>
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plasma is used. Milder reducing agents, such as sodium sulphite which do not include haem iron in the result, require heating at 90 to 95°C for about an hour to obtain a quantitative recovery of iron. Such a procedure does not recommend itself for inclusion in a rapid technique. The use of solid dithionite is perfectly satisfactory since a good grade product contains as little as 0.3 μg. Fe/g.; the addition of 100 mg. solid, five times the recommended amount, would increase the iron content of the final test solution by less than 1%. The addition of 0.05 ml. of a saturated solution (1 g. in 4 ml. water) of dithionite per determination is equally satisfactory, but the solution must be prepared immediately before use.

The proposed method compares favourably in accuracy with the method used by MacKenzie and Tindle, but the time for a single determination has been reduced by half. Recovery experiments were considered satisfactory.

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REFERENCES