The urinary excretion of assayable vitamin B\textsubscript{12} and radioactivity after parenteral \textsuperscript{58}Co B\textsubscript{12} in man

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SYNOPSIS  Evidence is presented that after injection of radioactive vitamin B\textsubscript{12} in man, there is a close correlation between the amount of radioactivity excreted and the amount of assayable vitamin B\textsubscript{12} excreted, and thus that the amount of radioactivity excreted is a true measure of the vitamin B\textsubscript{12} excreted. The possible reasons for this occurrence are discussed and it is suggested that in the body vitamin B\textsubscript{12} does not exist as such but as an analogue or active derivative.

It is generally assumed that, in man, vitamin B\textsubscript{12} (cyanocobalamin) is absorbed from the intestinal tract and functions in the body as an active substance without undergoing any chemical change. This view is based largely on the results of recovery experiments in which known amounts of vitamin B\textsubscript{12} are added to tissues; the increase in microbiological activity, as measured by assay methods considered to be specific for vitamin B\textsubscript{12}, is proportional to the amount of added vitamin B\textsubscript{12} and hence it seems likely that the original microbiological activity of the tissue is also due to vitamin B\textsubscript{12}.

If a patient who had no vitamin B\textsubscript{12} in his body and was not excreting any radioactive substance in his urine were given an injection of radioactive vitamin B\textsubscript{12} it would therefore be a reasonable assumption that the amount of radioactivity then found in the urine would be a true indication of the amount excreted via the kidney. For example, if this hypothetical patient were given 1,000 \mu g. radioactive vitamin B\textsubscript{12} intramuscularly and 20\% of the radioactivity was found in the urine then that 20\% would represent 200 \mu g. of vitamin B\textsubscript{12}. The normal subject, however, has between 1,500 and 5,000 \mu g. of non-radioactive vitamin B\textsubscript{12} in his body and the simple process outlined above cannot be assumed to occur because the possibility must be considered that the injected radioactive B\textsubscript{12} might equilibrate or mix with the non-radioactive body stores. Under these circumstances the amount of radioactivity appearing in the urine after an injection of radioactive vitamin B\textsubscript{12} would not be a true indication of the amount excreted but would represent only a fraction of that excreted. For example, if a patient with 3,000 \mu g. non-radioactive vitamin B\textsubscript{12} in his body were given 1,000 \mu g. radioactive vitamin B\textsubscript{12} intramuscularly and 20\% of the radioactivity was excreted in the urine then that 20\% would represent 200 \mu g. of vitamin B\textsubscript{12}.

Preliminary investigations showed that the normal subject excreted about 70\% of the radioactivity of an injection of 1,000 \mu g. radioactive B\textsubscript{12}; it was obvious that complete equilibration could not have taken place otherwise this would have represented about 2,800 \mu g. if the body stores were 3,000 \mu g. It was clearly important to determine whether any significant degree of equilibration did in fact occur and this could be detected by injecting radioactive vitamin B\textsubscript{12} and comparing the amount excreted in the urine as measured by the radioactivity present and by microbiological assay. If there was agreement with the two methods this would indicate that equilibration had not taken place whereas if equilibration did occur the amount of assayable vitamin B\textsubscript{12} in the urine would be greatly in excess of the amount indicated by the radioactivity present. For example, if a subject were given 1,000 \mu g. radioactive vitamin B\textsubscript{12} and 50\% of the radioactivity were excreted and a total of 500 \mu g. of vitamin B\textsubscript{12} was found by assay then this would be evidence against equilibration of significant degree: if, however, 50\% of the radioactivity were excreted and 750 \mu g. of vitamin B\textsubscript{12} were found by assay this would suggest that the 1,000 \mu g. of radioactive vitamin B\textsubscript{12} had equilibrated with 500 \mu g. of the body stores of non-radioactive vitamin B\textsubscript{12}.

Such an investigation supposes that the assay is accurate and fails to take into account the possibility of a constant error which would cause misleading
results. This error can be detected by duplicating the experiment in patients who have no body stores of vitamin B<sub>12</sub>. For example, if a normal subject were given 1,000 µg. radioactive vitamin B<sub>12</sub> intramuscularly and 50% of the radioactivity was excreted in the urine and 600 µg of vitamin B<sub>12</sub> was found by assay this would suggest either that there had been some equilibration or that the assay result was too high due to a constant error. If the same results were obtained in a patient with no body stores of vitamin B<sub>12</sub> then it would be clear evidence that the assay was giving a high result. Repeated injections of radioactive vitamin B<sub>12</sub> can be given to such control patients without prejudicing the results. If equilibration did occur then the injected radioactive vitamin B<sub>12</sub> would equilibrate with that remaining in the body from the previous injection. The use of control subjects would not only serve to reveal a constant error in the assay if such were present but would also reveal the degree of any equilibration which took place.

If it could be shown that equilibration of injected radioactive vitamin B<sub>12</sub> and non-radioactive body stores of vitamin B<sub>12</sub> did not occur then this finding would be of both practical and academic importance. Such a conclusion is drawn from the studies reported here.

**PLAN OF THE EXPERIMENT**

The object of the experiments was to compare the urinary excretion of radioactivity and of total assayable vitamin B<sub>12</sub> after parenteral injection of <sup>58</sup>Co vitamin B<sub>12</sub> in normal subjects and control subjects who were suffering from vitamin B<sub>12</sub> deficiency. To cover the usual range of doses used in clinical practice three different doses of <sup>58</sup>Co vitamin B<sub>12</sub> were used.

**PATIENTS, MATERIALS, AND METHODS**

The control patients were seven in-patients suffering from severe vitamin B<sub>12</sub> deficiency. Five were diagnosed as having Addisonian pernicious anaemia on the basis of a macrocytic anaemia, megaloblastic bone marrow, histamine-fast achlorhydria, low serum vitamin B<sub>12</sub> level, and a reticuloocyte response and rise in peripheral blood values after treatment with parenteral <sup>58</sup>Co vitamin B<sub>12</sub> only; after the study reported here the absorption of oral <sup>58</sup>Co vitamin B<sub>12</sub> was studied in these patients with and without intrinsic factor by the method of Schilling (1953) and the results supported the diagnosis. One patient was diagnosed as having adult coeliac disease on the basis of a macrocytic anaemia, megaloblastic bone marrow, free gastric acid, excess fat in the faeces, a low serum vitamin B<sub>12</sub> level, and reticuloocyte response and rise in peripheral blood values after treatment with <sup>58</sup>Co vitamin B<sub>12</sub> only; Schilling tests on this patient showed malabsorption of vitamin B<sub>12</sub> unaffected by the addition of intrinsic factor.

One patient came under observation with a macrocytic anaemia, megaloblastic bone marrow, and a low serum vitamin B<sub>12</sub> level four years after total gastrectomy. He also responded to parenteral <sup>58</sup>Co vitamin B<sub>12</sub> and Schilling tests showed that he failed to absorb it unless it was given with intrinsic factor. The control patients were treated by <sup>58</sup>Co vitamin B<sub>12</sub> intramuscularly daily.

The normal patients were in-patients with a variety of diseases in which depletion of body vitamin B<sub>12</sub> stores is not a known occurrence. All had normal serum vitamin B<sub>12</sub> levels. These patients were given a single intramuscular injection of <sup>58</sup>Co vitamin B<sub>12</sub>.

The nature and objects of the experiment were explained to the patients, all of whom cooperated willingly.

<sup>58</sup>Co vitamin B<sub>12</sub> was obtained from the Radiochemical Centre, Amersham, and diluted to a concentration of 0-5 µg., specific activity 0-5 µc., in 5 ml. normal saline, dispensed in dark glass rubber-capped bottles, autoclaved and stored at +4°C when not in use. The vitamin B<sub>12</sub> content was checked by microbiological assay every week the batch was in use. Ampoules of <sup>58</sup>Co vitamin B<sub>12</sub> for injection were prepared in the three dose ranges by adding 2 ml. of the <sup>58</sup>Co vitamin B<sub>12</sub> solution to 1,140 and 540 µg. non-radioactive vitamin B<sub>12</sub> and 3 ml. of the <sup>58</sup>Co vitamin B<sub>12</sub> solution to 54 µg. non-radioactive vitamin B<sub>12</sub>. The same tuberculin syringe was used for the preparation of these solutions and the preparation of the standard for isotope counting. Intramuscular injections were made by standard methods.

Urine was collected directly into acid-washed dark glass bottles. Pretreatment 24-hour collections were made on all patients. The total volume was noted and aliquots taken for isotope counting and microbiological assay. Samples for assay were diluted as required with vitamin B<sub>12</sub>-free water and stored at −20°C. till used: assays were performed as soon as possible.

**ISOTOPIC METHODS**

The radioactivity in 450 ml. aliquots of urine contained in plastic bottles was measured by the end-on method in a suitably modified Ecko scintillation counter type N 550 with thallium-activated sodium iodide crystal, 3-81 cm. diameter and 2-54 cm. deep, shielded by 3 in. of lead, and an Ecko automatic scaler type N 530A. A standard containing 0-2 µc. <sup>58</sup>Co vitamin B<sub>12</sub> and 2,000 µg. vitamin B<sub>12</sub> as a carrier in 450 ml. water was used. All samples were counted until 10,000 counts had been recorded in the 1,140 µg. and 540 µg. series, or until 5,000 counts had been recorded in the 54 µg. series.

**MICROBIOLOGICAL ASSAYS**

were conducted by a modification of the methods of Ross (1952) and Hutner, Bach, and Ross (1956) using Euglena gracilis 3 strain as the test organism. Samples were assayed in at least two different assay batches.

**STATISTICAL METHODS**

The results were analysed by standard statistical methods.

**RESULTS**

No radioactivity and only a negligible amount of assayable vitamin B<sub>12</sub> was found in the pretreatment...
collections in normal and in control patients. The results of the microbiological assays expressed as total assayable vitamin B$_{12}$ excreted in microgrammes and the amount of radioactivity excreted expressed as a percentage of the dose together with the regression line and 95% confidence limits for the control patients are plotted in Figs. 1, 2, and 3. The essential statistical results are set out in Table I.

It will be observed that the mean results for excretion in the control and normal patients do not differ greatly. This aspect will be discussed in a further communication.

DISCUSSION

A clear implication of the results is that, for practical purposes, parenteral radioactive vitamin B$_{12}$, in the doses used, does not equilibrate with the body stores of non-radioactive vitamin B$_{12}$, at least in the 24 hours after injection. In the 54 $\mu$g. series it is permissible for the regression lines and standard deviations to be drawn.
FIG. 3. Correlation of assayable vitamin $B_{12}$ and radioactivity excreted after injections of 1,140 $\mu$g. $^{58}$Co vitamin $B_{12}$ in normal and control subjects.

### TABLE I

ESSENTIAL STATISTICAL RESULTS OBTAINED IN THE INVESTIGATION

<table>
<thead>
<tr>
<th>Series (µg.)</th>
<th>Patients</th>
<th>No. of Readings</th>
<th>Mean Values</th>
<th>Coefficient of Correlation ($r$)</th>
<th>Regression of Isotope (x) on Assay (y)</th>
<th>S.E. of Estimate (%)</th>
</tr>
</thead>
<tbody>
<tr>
<td>54</td>
<td>Controls</td>
<td>30</td>
<td>683</td>
<td>1706</td>
<td>$x = 1.8809y + 4.2131$</td>
<td>2.256</td>
</tr>
<tr>
<td>54</td>
<td>Normals</td>
<td>18</td>
<td>5828</td>
<td>1412</td>
<td>$x = 1.9274y + 2.8259$</td>
<td>2.078</td>
</tr>
<tr>
<td>54</td>
<td>Controls</td>
<td>33</td>
<td>39236</td>
<td>7127</td>
<td>$x = 0.12731y + 21.314$</td>
<td>3.70</td>
</tr>
<tr>
<td>540</td>
<td>Normals</td>
<td>14</td>
<td>37300</td>
<td>6472</td>
<td>$x = 0.15675y + 6.2541$</td>
<td>1.88</td>
</tr>
<tr>
<td>1,140</td>
<td>Controls</td>
<td>30</td>
<td>82597</td>
<td>6663</td>
<td>$x = 0.05937y + 17.595$</td>
<td>6.91</td>
</tr>
<tr>
<td>1,140</td>
<td>Normals</td>
<td>15</td>
<td>82788</td>
<td>6997</td>
<td>$x = 0.05696ly + 22.807$</td>
<td>5.079</td>
</tr>
</tbody>
</table>
tion curves to be extended to meet the abscissa and ordinate. The lower SD curve for the control group (Fig. 1) lies on the point abscissa O, ordinate O, while for the normal group the lower SD curve meets the ordinate at the point abscissa O ordinate -1·88, the regression line for the controls meets the ordinate at the point abscissa O ordinate 2·83 and the upper SD curve for the controls meets the ordinate at the point abscissa O ordinate 7·54. These findings are good evidence that the lack of equilibration occurs in the entire dose range up to 1,140 μg.

It is difficult to explain the phenomenon on the basis of the concept that vitamin B₁₂ is present as such in man. An obvious explanation would be that the injected radioactive vitamin B₁₂ is excreted so rapidly that there is insufficient time for equilibration to occur; detailed observations on the rate of excretion of injected ⁵⁸Co vitamin B₁₂ suggest that this explanation is unlikely to be wholly satisfactory.

The explanation may be related to the fact that the preponderant mass of vitamin B₁₂ in the body is tissue bound. The process of binding involves at least physical changes, since the assayable vitamin B₁₂ in liver is microbiologically active but is not dialysable, and it is pertinent to consider whether the process may not also involve chemical changes in the structure of the vitamin B₁₂ molecule, i.e., whether vitamin B₁₂ absorbed from dietary sources is converted to an analogue in the process of binding to body tissues. If this were the case one would not expect the injected radioactive vitamin B₁₂ to equilibrate with the analogue until it also has been converted by the process of binding, by which time it would be fixed to the tissues and no longer free to be excreted. The proportion which is excreted would be that which the tissues have been unable to bind in the time available.

Recently biochemical evidence has been produced which suggests that a very considerable proportion of the cobalamins in rabbit liver are present not as vitamin B₁₂ but as coenzymes which may be the metabolically active form of vitamin B₁₂. There is also evidence that these coenzymes can be converted to vitamin B₁₂ by procedures commonly used in microbiological assays of tissues such as exposure to heat, light, or to cyanide ions (Weissbach, Toohy, and Barker, 1959). This suggests that in microbiological assays the substance which is actually measured is vitamin B₁₂ derived from body stores of the microbiologically active form. In addition, one of these coenzymes, 5, 6-dimethylbenzimidazolylcobamide coenzyme, has been shown to be as metabolically active in patients with pernicious anaemia as vitamin B₁₂ at the same dosage level (Wasserman, Estren, Brody, and Herbert, 1960).

It is also clear from the results that when man is given such injections of radioactive vitamin B₁₂ the radioactivity found in the urine in the subsequent 24 hours is a true indication of the amount of assayable vitamin B₁₂ excreted and that this occurs regardless of its body stores. This finding is of considerable practical importance for it means that the isotope method can be used with confidence instead of the microbiological assay in determining the total urinary loss of vitamin B₁₂ after an injection of radioactive vitamin B₁₂, at least for the doses used. The isotope method, which makes use of a very small dose of radioactivity, has obvious advantages over the assay method; as a technical method it is inherently more accurate and reliable results may be obtained quickly and with little expenditure of time or labour. The practical application of this method will be dealt with in subsequent papers.

I am grateful to Dr. J. A. W. McCluskie for his advice and criticism and to Professor E. J. Wayne for facilities in his department. It is a pleasure to thank Dr. R. A. Robb for his advice on the statistical problems and Dr. M. M. Bluhm for his advice on the isotopic aspects. The ⁵⁸Co vitamin B₁₂ was supplied by the Regional Physics Department, Western Regional Hospital Board.

REFERENCES